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mole of water is the water molecule, the representative particle in a mole of copper is the copper atom, and the representative particle in a mole of sodium chloride is the formula unit. 310 Chapter 11 The Mole Figure 11-2 The amount of each substance shown is 6.02 $\times 10^{23}$ or one mole of representative particles. The representative particle for

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The mole The mole, abbreviated mol, is the SI base unit used to measure the amount of a substance. A mole is defined as the number of carbon atoms in exactly 12 g of pure carbon-12. Through years of experimentation, it has been established that a mole of anything contains

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Therefore its molar mass is .2 Mass and the Mole Molar mass The mass in grams of a mole of any pure substance is called its molar mass. Section 11.94 amu. Ex. An atom of manganese has the atomic mass of 54.

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b. 1 mole of C₂H₆ 10. Arrange the following from the smallest number of representative particles (top = Item 1) to the largest number of representative particles (Item 3).

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Mass and the Mole (11.2 pgs. 313-319) The relative scale of atomic masses uses the isotope carbon-12 as the standard. Each atom of carbon-12 has a mass of 12 atomic mass units (amu).

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A mole ratio is a ratio between the numbers of moles of any two of the substances in a balanced chemical equation. For example, consider the reaction shown in Figure 11.2.

Chapter 11: Stoichiometry

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Episode 11 €“ The Mole Answer Key 1. Why is it important to use the correct amount of materials in a chemical reaction? If too little is used the reaction may not proceed very far. The use of too much chemical

may result in waste. 2. What names are given to the materials at the beginning and end of a chemical reaction? Reactants and products. 3.

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Atomic mass unit and the mole "amu definition: $12\text{C} = 12\text{ amu}$. " The atomic mass unit is defined this way. " $1\text{ amu} = 1.6605 \times 10^{-24}\text{ g}$ " How many 12C atoms weigh 12 g ? " 6.02×10^{23} 12C weigh 12 g . " Avogadro's number

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Mass in grams of one mole of any pure substance. (g/mol) Equal to the atomic mass.

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Chapter 11 - Supplemental Problems - The Mole. 1. Identify and calculate the number of representative particle in each of the following quantities. a. $2.14\text{ moles of gold}$ a. 6.35 g lithium b. $0.151\text{ mole of nitrogen oxide (NO)}$ b. 346 g zinc c. $11.5\text{ moles of potassium bromide (KBr)}$ c. 115 g nickel 2.

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