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5. How much (grams) barium chloride is necessary to react with the silver nitrate in Problem 4? Chemistry IF8766 Page 64 Instructional Fair, Inc.

Stoichiometry: Mass-Mass Problems

b. mass - volume and volume - volume problems Since chemical equations for chemical reactions state the relative numbers of moles for each reactant and product, these numbers can tell us the volumes at STP for

those substances that

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Stoichiometry: Mixed Problems (KEY) 1) $N_2 + 3H_2 \rightarrow 2NH_3$ What volume of NH_3 at STP is produced if 25.0 of N_2 is reacted with an excess of H_2 ? 3 3 3 2 3 2 2 2 40.0L NH_3 1mol NH_3 22.4L NH_3 1mol N_2 2mol NH_3 28.0g N_2 25.0g N_2 1mol N_2 $\frac{2}{1} = 2$) $2KClO_3 \rightarrow 2KCl + 3O_2$ If 5.0g of $KClO_3$ is decomposed, what volume of O_2 is produced at STP? 2

Stoichiometry: Mixed Problems (KEY)

Chemistry IF8766 62 . STOICHIOMETRY: MASS-MASS PROBLEMS 1. $2KClO_3 \rightarrow 2KCl + 3O_2$ Name How many grams of potassium chloride are produced if 25 g of potassium chlorate decompose? / 25 ken How many grams of hydrogen are necessary to react completely with 50.0 g of nitrogen in the above reaction?

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How many grams of ammonia are produced in the reaction in Problem 2 4. $2AgNO_3 + BaCl_2 \rightarrow 2AgCl + Ba ...$ (grams) barium chloride is necessary to react with the silver nitrate in Problem 4? Chemistry IF8766 Page 64 Instructional Fair, Inc. Title: Microsoft Word - pg 64 - Stoichiometry - mass-mass problems.doc Author: Nathan Created Date:

Stoichiometry: Mass-Mass Problems

Name: _____ Moles and Mass Directions: Determine the number of moles in each of the quantities below. 1) 52 g of $NaCl$ ANSWER: 0.89 mole 2) 145 g of H_2SO_4 ANSWER: 1.48 mole 3) 110. g of $KMnO_4$

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Name _ X 0 (DETERMINING- MOLECULAR FORMULAS (TRUE FORMULAS) Solve the problems below. 1. The empirical formula of a compound is NO2. Its molecular mass is 92 g/mol.

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