

## DOWNLOAD COLLIGATIVE PROPERTIES GIZMO ANSWERS

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Colligative properties are those properties of solutions that depend on the number of dissolved particles in solution, but not on the identities of the solutes. Molarity (M), molality (m) and the colligative properties

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Determine how the physical properties of a solvent are dependent on the number of solute particles present. Measure the vapor pressure, boiling point, freezing point, and osmotic pressure of pure water and a variety of solutions.

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Gizmo Warm-up Adding salt or other substances to water can affect the temperature at which it freezes or boils. These effects and others, known as colligative properties, are explored in the Colligative Properties Gizmo. Check that No solute is selected and the Air temp. is 25 °C. 1. Describe the motion of the water molecules in the beaker.

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Colligative Properties Problems - Answers Remember that Colligative properties depend only on the number or concentration of particles in a solution. The properties are, for ideal solutions, independent of the kind or size of the particles, whether ionic, large or small etc.

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### **Colligative Properties Gizmo Worksheet Answer Key**

Colligative comes from colligate "to tie together Colligative properties have common origin Colligative properties depend on amount of solute but do not depend on its chemical identity ...

### **Colligative Properties - College of DuPage**

Colligative Properties Worksheet " Answers 1) If I add 45 grams of sodium chloride to 500 grams of water, what will the melting and boiling points be of the resulting solution?

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Colligative Properties. Determine how the physical properties of a solvent are dependent on the number of solute particles present. Measure the vapor pressure, boiling point, freezing point, and osmotic pressure of pure water and a variety of solutions.

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### **ColligativePropertiesSE - Name Kristopher Ward Student**

WORKSHEET:SOLUTIONS AND COLLIGATIVE PROPERTIES SET A: 1. Find the molarity of all ions in a solution that contains 0.165 moles of aluminum chloride in 820. ml solution. Answer:  $[Al^{3+}] = 0.201 M$  ,  $[Cl^-] = 0.603M$ . 2. Find the molarity of each ion present after mixing 27 ml of 0.25 M  $HNO_3$  with 36 ml of 0.42 M  $Ca(NO_3)_2$  ... Microsoft Word ...

### **WORKSHEET:SOLUTIONS AND COLLIGATIVE PROPERTIES SET A**

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Lab 9. Colligative Properties ... Answer the pre-lab questions on the last 2 pages, to be turned in at the start of the lab period. Purpose In this online lab activity you will use freezing point depression and boiling point elevation data to calculate the  $K_f$  and  $K_b$  of water and explore the van't Hoff factor for electrolytes.

### **Lab 9. Colligative Properties an Online Lab Activity**

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